

★ How are formula names of ionic compounds written?

Chemical formula: a group of symbols that shows the ratio of elements in a compound.

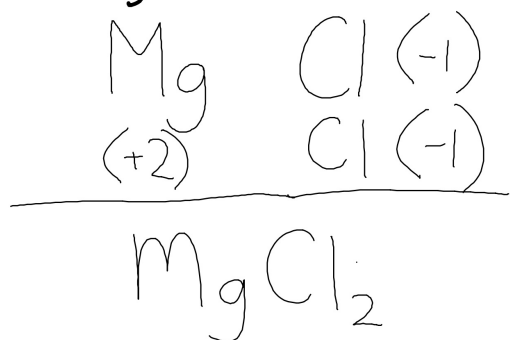
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↳ For Ionic Compounds.....

- 1) always write the symbol for the positive ion first!
- 2) then, you write the negative ion next
- 3) add subscripts as needed to balance the charges.

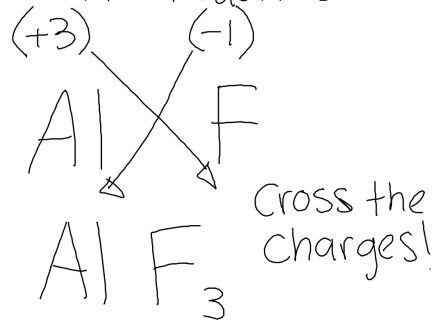
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ex) Magnesium Chloride



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ex) Aluminum Fluoride



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Naming Ionic Compounds

- 1) Name the 1st ion (+) as it appears on the P.T.
- 2) a) end the name in "-ide" if the second ion is a single element
b) end name in "-ate" or "-ite" for polyatomic ions.

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ex) MgO

Magnesium Oxide
(single element)

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ex) KNO_3 (pg. 54)

Potassium Nitrate

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Properties of Ionic Compounds

- ↳ these form hard, brittle,
① crystals (NaCl)
- ② High melting points
- ③ ability to conduct electricity
when dissolved b/c of charged
particles

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